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Key indicators

Single-crystal X-ray study
 $T = 293\text{ K}$
Mean $\sigma(\text{Ga-S}) = 0.006\text{ \AA}$
 R factor = 0.053
 wR factor = 0.186
Data-to-parameter ratio = 19.7

For details of how these key indicators were automatically derived from the article, see <http://journals.iucr.org/e>.

Rubidium thiogallate

RbGaS_2 , for which only cell parameters were previously known, consists of GaS_2 layers interleaved with Rb atoms.

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Comment

There are eight $AG\text{aS}_2$ (A = alkali metal, Tl, Cu and Ag) compounds known. They contain gallium in tetrahedral coordination and have four types of diverse structures, ranging from one-dimensional to three-dimensional. For RbGaS_2 and TlGaS_2 , only unit-cell parameters are available, whereas for the five $AG\text{aS}_2$ (A = Li, K, Cs, Ag Cu) compounds, structures have been determined by single-crystal X-ray diffraction. LiGaS_2 is derived from zincite, wherein Li and Ga atoms occupy the tetrahedral sites in an hcp array of S atoms in an ordered fashion (Leal-Gonzalez *et al.*, 1990). The structure of KGaS_2 was solved in the non-centrosymmetric space group Aa (Lemoine *et al.*, 1984) but is actually the same as that of layered TlGaSe_2 , with the centrosymmetric space group $C2/c$ (Müller & Hahn, 1978; Henkel *et al.*, 1982). NaGaS_2 , RbGaS_2 and TlGaS_2 are believed to have the same TlGaSe_2 structure

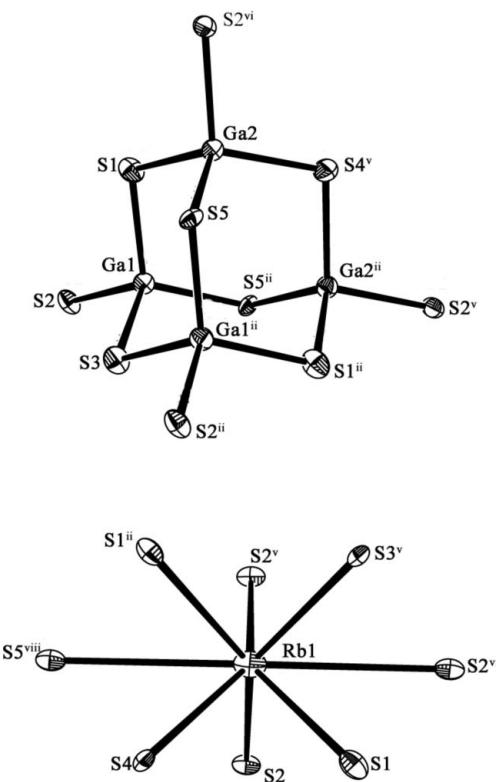
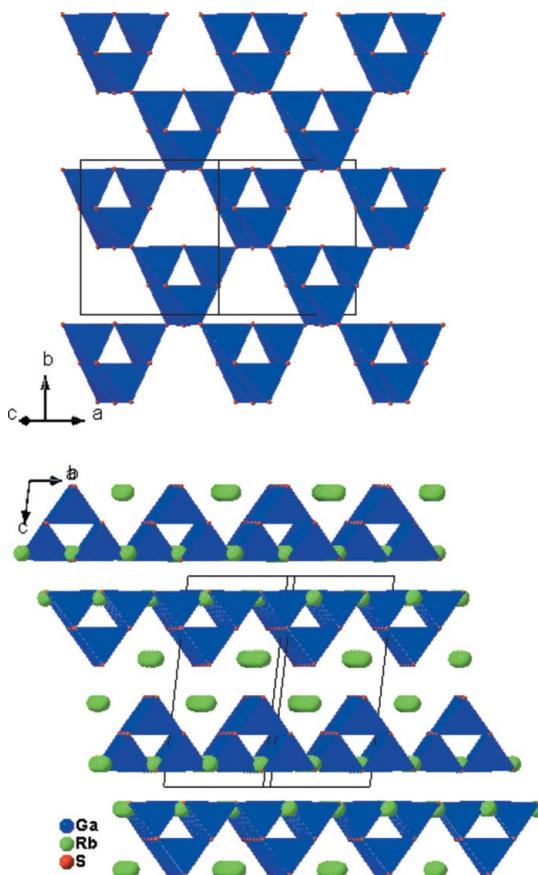


Figure 1

ORTEP-3 (Farrugia, 1997) diagrams of the tetrahedral Ga_4S_{10} unit (top) and bicapped trigonal prism of Rb_1S_8 (bottom). Displacement ellipsoids are drawn at the 50% probability level. [Symmetry codes as in Table 1.]

**Figure 2**

Polyhedral representation of the GaS₂ layer viewed along [101] (top), and the unit cell contents of RbGaS₂ viewed along the *a* axis (bottom).

(Hoppe, 1965; Schubert & Hoppe, 1970; Klepp, 1992; Müller *et al.*, 1974). CsGaS₂ is a one-dimensional compound, in which a GaS₄ tetrahedron shares edges with two other tetrahedra to form GaS₂ chains that are separated by Cs atoms (Schmitz & Bronger, 1975). CuGaS₂ and AgGaS₂ compounds are of chalcopyrite type (Abrahams & Bernstein, 1973*a,b*). We report here the single-crystal X-ray structure of RbGaS₂, confirming that it is isotopic with TlGaSe₂. It is noteworthy that the crystal structure elucidation of RbGaS₂ has so far remained elusive, because of twinning and stacking faults that are common for crystals of such layered compounds (Kienle *et al.*, 2004). KInS₂ (Lowe-Ma *et al.*, 1991; Eisenmann & Hofmann, 1991*a*), KInSe₂ (Krebs, 1983), RbInS₂ (Müller & Hahn, 1978), CsInS₂, KTlS₂, RbTlS₂, CsTlS₂ (Schubert & Hoppe, 1970), TlAlSe₂, TlAlS₂, TlInS₂ (Müller *et al.*, 1974) and NaAlSe₂ (Eisenmann & Hofmann, 1991*b*) are the other known compounds with the TlGaSe₂ structure.

In RbGaS₂, four GaS₄ tetrahedra form adamantane-like Ga₄S₁₀ units (Fig. 1). These units are connected by corners to one another to give rise to GaS₂ layers, which are stacked along the *c* axis and interleaved with Rb atoms (Fig. 2). The Rb atoms have bicapped trigonal prismatic coordination, as shown for Rb1 in Fig. 1. Bond lengths and angles compare well with those reported in the literature (Devi & Vidyasagar, 2002; Schmitz & Bronger, 1975; Lemoine *et al.*, 1984; Yao & Ibers, 2004).

Experimental

The chemicals used for synthesis were of greater than 99% purity and were purchased from CERAC Inc. and Alfa Aesar. A mixture of Rb₂CO₃ (0.2 g, 0.086 mmol), Ga₂S₃ (0.0139 g, 0.059 mmol) and S (0.1388 g, 4.32 mmol) was sealed in an evacuated fused-silica tube of 13 cm length and 1.3 cm diameter, heated in a furnace at 1023 K for 4 d and then cooled to room temperature over a period of 4 d. The product was washed with water to dissolve away the rubidium polysulfide flux, enabling the isolation of single-phase RbGaS₂ in the form of plate-like red–brown crystals (0.0608 g, 75% based on Ga₂S₃). The powder X-ray diffraction pattern of the sample compares well with that simulated using the program LAZY PULVERIX (Yvon *et al.*, 1977) on the basis of the single-crystal X-ray structure.

Crystal data

RbGaS ₂	$D_x = 3.499 \text{ Mg m}^{-3}$
$M_r = 219.31$	Mo $K\alpha$ radiation
Monoclinic, $C2/c$	Cell parameters from 25 reflections
$a = 10.484 (15) \text{ \AA}$	$\theta = 10\text{--}15^\circ$
$b = 10.468 (15) \text{ \AA}$	$\mu = 18.99 \text{ mm}^{-1}$
$c = 15.390 (16) \text{ \AA}$	$T = 293 (2) \text{ K}$
$\beta = 99.69 (10)^\circ$	Block, brown
$V = 1665 (4) \text{ \AA}^3$	$0.2 \times 0.1 \times 0.1 \text{ mm}$
$Z = 16$	

Data collection

Enraf-Nonius CAD-4 diffractometer	$R_{\text{int}} = 0.077$
ω -2 θ scans	$\theta_{\text{max}} = 25.0^\circ$
Absorption correction: ψ scan (North <i>et al.</i> , 1968)	$h = -12 \rightarrow 8$
$T_{\text{min}} = 0.076$, $T_{\text{max}} = 0.150$	$k = -12 \rightarrow 8$
3083 measured reflections	$l = -18 \rightarrow 18$
1476 independent reflections	2 standard reflections
895 reflections with $I > 2\sigma(I)$	frequency: 60 min
	intensity decay: none

Refinement

Refinement on F^2	$w = 1/[\sigma^2(F_o^2) + (0.0817P)^2 + 20.9393P]$
$R[F^2 > 2\sigma(F^2)] = 0.053$	where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.186$	$(\Delta/\sigma)_{\text{max}} < 0.001$
$S = 1.19$	$\Delta\rho_{\text{max}} = 1.32 \text{ e \AA}^{-3}$
1476 reflections	$\Delta\rho_{\text{min}} = -1.35 \text{ e \AA}^{-3}$
75 parameters	Extinction correction: SHELXL97
	Extinction coefficient: 0.00029 (8)

Table 1
Selected interatomic distances (\AA).

Ga1–S1	2.263 (5)	Rb1–S2 ^{vi}	3.380 (6)
Ga1–S3 ⁱ	2.268 (6)	Rb1–S1 ^{vii}	3.384 (6)
Ga1–S5 ⁱ	2.294 (5)	Rb1–S2	3.717 (7)
Ga1–S2	2.300 (5)	Rb1–S2 ^v	3.734 (7)
Ga2–S1	2.263 (5)	Rb2–S2 ^{viii}	3.318 (6)
Ga2–S4 ⁱⁱ	2.271 (6)	Rb2–S2 ^{ix}	3.350 (6)
Ga2–S5	2.288 (5)	Rb2–S1 ^x	3.419 (6)
Ga2–S2 ⁱⁱⁱ	2.296 (5)	Rb2–S1	3.424 (6)
Rb1–S5 ^{iv}	3.330 (6)	Rb2–S3 ⁱ	3.425 (5)
Rb1–S3 ^v	3.365 (7)	Rb2–S4	3.425 (4)
Rb1–S4	3.366 (7)	Rb2–S5 ^{xi}	3.728 (7)
Rb1–S1	3.370 (5)	Rb2–S5	3.731 (7)

Symmetry codes: (i) $-x, y, -z + \frac{1}{2}$; (ii) $x - \frac{1}{2}, y - \frac{1}{2}, z$; (iii) $-x + \frac{1}{2}, y - \frac{1}{2}, -z + \frac{1}{2}$; (iv) $x + \frac{1}{2}, -y + \frac{1}{2}, z - \frac{1}{2}$; (v) $x + \frac{1}{2}, y - \frac{1}{2}, z$; (vi) $-x + \frac{1}{2}, -y + \frac{1}{2}, -z$; (vii) $-x + 1, y, -z + \frac{1}{2}$; (viii) $-x + \frac{1}{2}, -y + \frac{1}{2}, -z + 1$; (ix) $x, -y + 1, z + \frac{1}{2}$; (x) $-x + \frac{1}{2}, y + \frac{1}{2}, -z + \frac{1}{2}$; (xi) $x + \frac{1}{2}, y + \frac{1}{2}, z$.

The highest peak and the deepest hole in the final Fourier map are located 0.85 \AA from Ga2 and 1.18 \AA from Ga1, respectively.

Data collection: *CAD-4 Software* (Enraf–Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *XCAD4* in *WinGX* (Farrugia, 1999); program(s) used to solve structure: *SHELXL86* (Sheldrick, 1986); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *DIAMOND* (Brandenburg, 1999) and *ORTEP-3* (Farrugia, 1997); software used to prepare material for publication: *SHELXL97*.

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References

- Abrahams, S. C. & Bernstein, J. L. (1973a). *J. Chem. Phys.* **59**, 1625–1629.
 Abrahams, S. C. & Bernstein, J. L. (1973b). *J. Chem. Phys.* **59**, 5415–5422.
 Brandenburg, K. (1999). *DIAMOND*. Crystal Impact GbR, Bonn, Germany.
 Devi, M. S. & Vidyasagar, K. (2002). *J. Chem. Soc. Dalton Trans.* pp. 4751–4754.
 Eisenmann, B. & Hofmann, A. (1991a). *Z. Kristallogr.* **195**, 318–319.
 Eisenmann, B. & Hofmann, A. (1991b). *Z. Kristallogr.* **197**, 171–172.
 Enraf–Nonius (1989). *CAD-4 Software*. Version 5.0. Enraf–Nonius, Delft, The Netherlands.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
 Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
 Henkel, W., Hochheimer, H. D., Carbone, C., Werner, A., Ves, S. & von Schnering, H. G. (1982). *Phys. Rev. B*, **26**, 3211–3221.
 Hoppe, R. (1965). *Bull. Soc. Chim. Fr.* pp. 1115–1121.
 Kienle, L., Duppel, V., Simon, A., Schlosser, M. & Jarchow, O. (2004). *J. Solid State Chem.* **177**, 6–16.
 Klepp, K. O. (1992). *Z. Naturforsch. Teil B*, **47**, 937–941.
 Krebs, B. (1983). *Angew. Chem. Int. Ed. Engl.* **22**, 113–134.
 Leal-Gonzalez, J., Melibary, S. A. & Smith, A. J. (1990). *Acta Cryst. C* **46**, 2017–2019.
 Lemoine, P., Carré, D. & Guittard, M. (1984). *Acta Cryst. C* **40**, 910–912.
 Lowe-Ma, C. K., Kipp, D. O. & Vanderah, T. A. (1991). *J. Solid State Chem.* **92**, 520–530.
 Müller, D. & Hahn, H. (1978). *Z. Anorg. Allg. Chem.* **438**, 258–272.
 Müller, D., Poltmann, F. E. & Hahn, H. (1974). *Z. Naturforsch. Teil B*, **29**, 117–118.
 North, A. C. T., Phillips, D. C. & Mathews, F. S. (1968). *Acta Cryst. A* **24**, 351–359.
 Schmitz, D. & Bronger, W. (1975). *Z. Naturforsch. Teil B*, **30**, 491–493.
 Schubert, H. & Hoppe, R. (1970). *Z. Naturforsch. Teil B*, **25**, 886–887.
 Sheldrick, G. M. (1986). *SHELXS86*. University of Göttingen, Germany.
 Sheldrick, G. M. (1997). *SHELXL97*. University of Göttingen, Germany.
 Yao, J. & Ibers, J. A. (2004). *Acta Cryst. E* **60**, i95–i96.
 Yvon, K., Jeitschko, W. & Parthé, E. (1977). *J. Appl. Cryst.* **10**, 73–74.